

## Naming:

1. ammonium oxide
2. sodium sulfate
3. calcium hydroxide
4. Calcium carbonate
5. lead (II) carbonate

# Formulas





# How do you write the names?

- We use the same procedure as ionic compounds except we have to include prefixes in the name to indicate how many atoms of each there are.

**Prefix indicates the number of atoms of each element:**

- mono - 1
- di - 2
- tri - 3
- tetra - 4
- penta - 5
- hexa - 6

•\*\*\* If the first element only has one atom we do not have to place the "mono".

# Molecular (covalent) Compounds: Examples

## Examples:

1. NO

Nitrogen monoxide

➤ 2. N<sub>2</sub>O

dinitrogen monoxide

➤ 3. N<sub>2</sub>O<sub>5</sub>

dinitrogen pentoxide

➤ 4. CO<sub>2</sub>

Carbon dioxide

# Practice:

Write the names:

- $\text{NO}_3$  Nitrogen trioxide
- $\text{SO}_2$  Sulfur dioxide
- $\text{CO}$  Carbon monoxide
- $\text{SiBr}_2$  Silicon dibromide
- $\text{PCl}_5$  Phosphorus pentachloride
- $\text{SO}_3$  Sulfur trioxide
- $\text{SO}_4$  Sulfur tetraoxide

# How do you write the formulas for molecular compounds?

Write the symbol of the element situated the furthest to the left in the periodic table first, then the second non-metal symbol.

- Use the prefixes to determine the formula.

# Examples

## ➤ Write the Formulas

- nitrogen tri-bromide



- nitrogen tetra-iodide



- nitrogen dioxide



- sulphur penta-oxide





# Practice!

1. carbon tetrachloride



2. silicon dioxide



3. dihydrogen monoxide



4. sulphur trioxide



5. di-nitrogen monoxide

